

作业 20091028

A

ACT: P133: 4.20, 4.21, 4.22

B

1. 思考: ACT p132:4.8
2. A 25.00-mL sample of unknown containing Fe^{3+} and Cu^{2+} required 16.06 mL of 0.05083 mol/L EDTA for complete titration. A 50.00-mL sample of the unknown was treated with NH_4F to protect the Fe^{3+} . Then the Cu^{2+} was reduced and masked by addition of thiourea (硫脲). Upon addition of 25.00-mL of 0.05083 mol/L EDTA, the Fe^{3+} was liberated from its fluoride complex and formed an EDTA complex. The excess EDTA required 19.77 mL of 0.01883 mol/L Pb^{2+} to reach an end point, using xylenol orange (二甲酚橙). Find the concentration of Cu^{2+} in the unknown.